Name: _____ Date: _____ Period: _____

Chemistry Vocabulary

_____ 1. group of unreactive elements

_____ 2. the most reactive group of metals, they're in group 1

_____ 3. what chemistry is all about, they're in the outer shell

_____ 4. There are approximately 100 -- listed on the periodic table.

_____ 5. elements to the left side of the periodic table

_____ 6. group 2 elements

_____ 7. a horizontal row in the periodic table

_____ 8. elements on the upper right side of the periodic table

_____ 9. a vertical column in the periodic table

_____10. reactive nonmetals in group 17

_____11. elements along either side of the zig-zag line in the periodic table

_____12. The smallest part of an element that can exist alone or in combination with others.

_____13. The ability to allow electricity and heat to pass through

_____14. can be pounded into sheets

_____15. an arrangement of elements first developed by Mendeleev

_____16. The particles that make up an atom; protons, neutrons, and electrons.

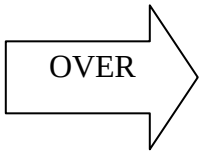
_____17. The number of protons and neutrons in an atom.

_____18. The average mass of all the isotopes of an element.

_____19. Different atoms of the same element with different number of neutrons.

_____20. elements in groups 3-12

_____21. can be drawn into wire



Fill out the chart.

		Mass Number	Atomic Number	Protons	Neutrons	Electrons	Hyphenated Notation
1	$^{12}_{6}C$						
2		75	33				
3				56	81		
4							Carbon-13

Which two numbers above represent isotopes?

Draw a Bohr Model and Lewis Dot Structure for the **oxygen** atom.

Bohr Model: Lewis Dot Structure:

How many valance electrons does oxygen have? _____

How many electrons would oxygen have to gain to reach a stable state? _____

How many electrons would oxygen have to lose to reach a noble state? _____