

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

	<b>Scientist</b>	<b>Experiment</b>
___ 1	Bohr	A. Cathode Ray
___ 2	Thompson	B. Gold foil
___ 3	Rutherford	C. Hydrogen's line spectrum

	<b>experiment</b>	<b>Evidence from experiment</b>
___ 4	Cathode ray	A. Positive particles were deflected by something positive
___ 5	Gold foil	B. Atoms emit only lines of color when energized which means
___ 6	Hydrogen's line spectrum	C. A ray is attracted to a positive plate

	<b>model</b>	<b>Conclusion from experiment</b>
___ 7	Nuclear	A. Electrons only exist at certain distances from the nucleus
___ 8	Plum Pudding	B. Atoms remain intact & are not destroyed in chemical reactions
___ 9	Planetary	C. There must be electrical charges associated with matter
___ 10	Billiard Ball	D. All of the positive matter in an atom must be small & centrally located

	<b>model</b>	<b>experiment</b>
___ 11	Nuclear	A. Cathode Ray
___ 12	Plum Pudding	B. Gold foil
___ 13	Planetary	C. Hydrogen's line spectrum

**List the 4 models of the atom in chronological order. Then describe the model & what distinguishes it from the previous model**

	<b>Model</b>	<b>Description/distinction</b>
1.	_____	- _____
2.	_____	- _____
3.	_____	- _____
4.	_____	- _____